

FORM 4 END TERM 2 2020
CHEMISTRY PAPER 2

1. Study the information given below and answer the questions that follow.

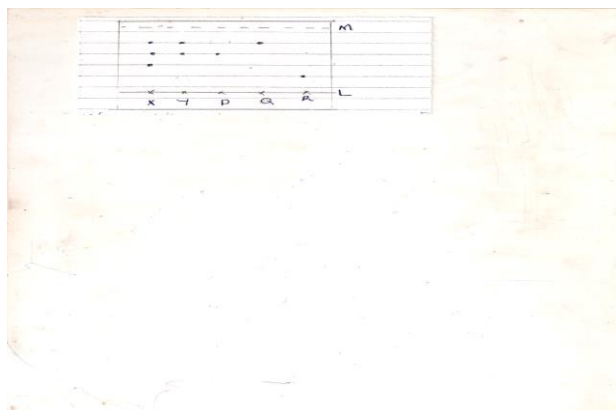
Element	Atomic radius(nm)	Ionic radius nm	Formula of oxide	Melting point(⁰ c)
A	0.364	0.421	A ₂ O	-119
D	0.830	0.711	D O ₂	837
E	0.592	0.485	E ₂ O ₃	1466
G	0.381	0.446	G ₂ O ₃	242
J	0.762	0.676	J O	1054

- a. Which elements are non-metals .Give a reason?(2mks)
- b. i) Write a formula of a compound formed when J combines with A(1mk)
- ii) What type of bond exist between J and D.(1mk)
- c. Explain why the melting point of the oxide of E is higher than that of the oxide of G.(2mks)
- d. i) Which two elements would react with each other most vigorously.Give a reason.(2mks)
- ii) Which element would be suitable for making utensils for boiling water.State two properties that make the elements suitable for the use.(2mks)

- e. Elements Q and R have electronic configuration 2.8.2 and 2.8.6. respectively.
 i) Explain why the ionic radius of R is expected to be greater than its atomic radius. (1mk)

ii) Write the equation for the reaction between Q and R. (1mk)

2. The chromatogram below is of and acid enzyme x and y and three simple sugar P, Q and R.



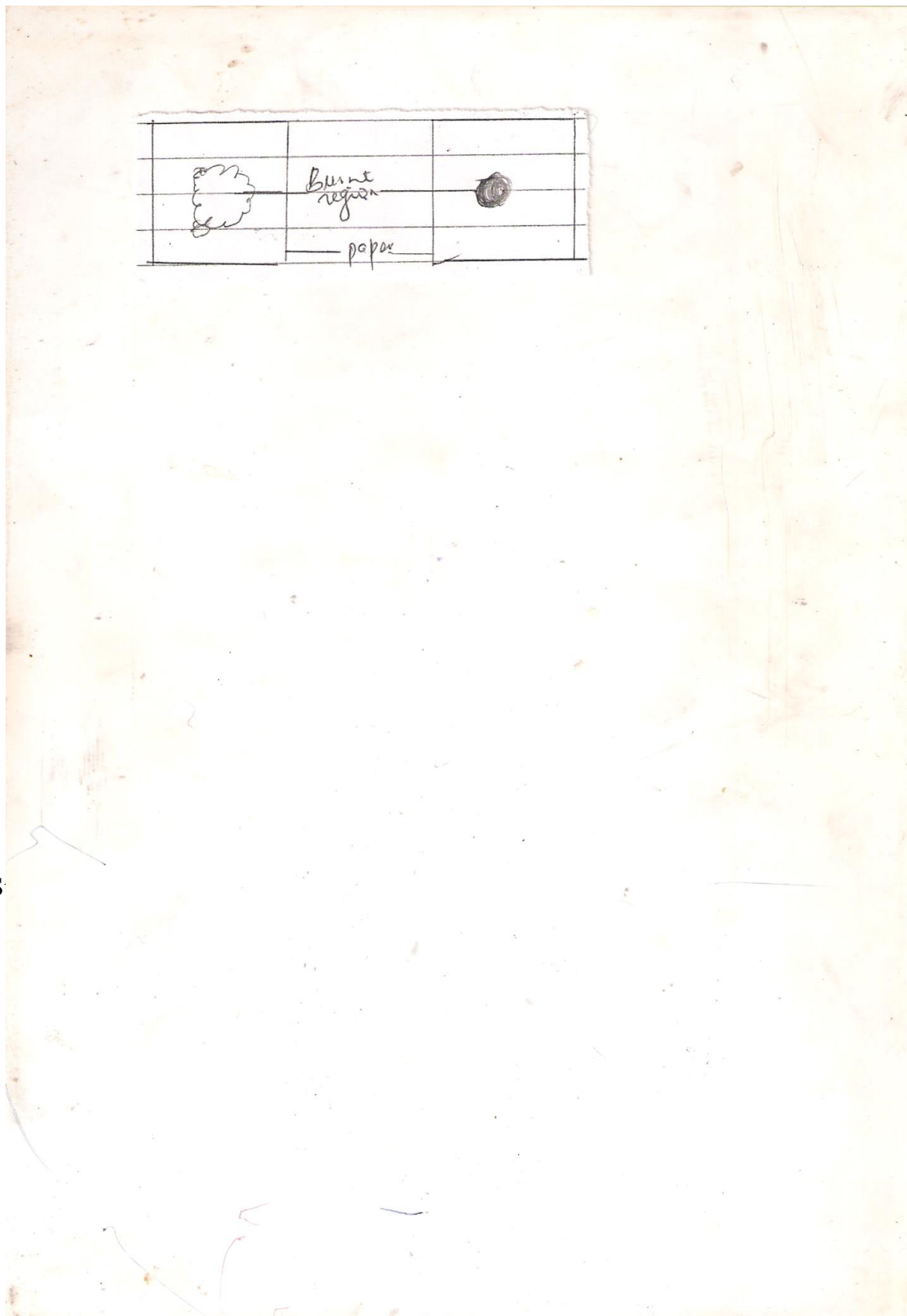
- a. i) Name two simple sugars present in both x and y. (2mks)
 ii) Name lines L and M. (2mks)

L-

M-

- iii) What property is exhibited by simple sugar x. (1mk)

b. Two pieces of paper were lowered into different Bunsen burner flames and removed quickly. The results were as shown below.



2mks)

3. S

18
om

Dilute hydrochloric acid

$\text{Pb}(\text{NO}_3)_2$

Excess

NaOH

a. Identify

i. White ppt I (1mk)

ii. Solution II (1mk)

iii. Residue II (1mk)

b. Write ionic equation for the reactions colourless solution (II) with $\text{Pb}(\text{NO}_3)_2$ 1mk

c. Write observations that would be made when ammonia solution is added drop wise till in excess to the colourless solution(II) 2mks

d. Below are P^{H} values of some solutions

Solution	Z	Y	X	W
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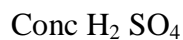
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P^H	6.5	3.5	2.2	7.2
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- i. Which solution is likely to be
 - a. Acidic rain (1mk)
 - b. Potassium hydroxide (1mk)
 - ii. A basic substance V reacted with both solutions Y and X. What is the nature of V. (2mks)
 - iii. Name two substances that show these characteristics in question (ii) above. (2mks)
4. A sample of crude oil was heated and its vapour passed over red-hot pumice stone. A mixture of gases was evolved which decolourised bromine in tetra chloromethane and burnt in air with a yellow flame.
- a. What process is taking place when the vapour from the crude oil passes over heated pumice. (1mk)
 - b. Name the most likely type of compound causing decolourisation of the bromine solution. (1mk)
 - c. Name two compounds formed when the gas mixture above burns in air. (1mk)

ii. Study the flow chart below and answer the questions that follow.



high

pressure

H₂

O₂(Excess)

Line water

Na

- a. Identify substances (4mks)
- A-
 - B-
 - F-
 - G-
- b. Write down the equation for the formation of
- i. Substance C
 - ii. E and F
 - iii. Gas G
- c. Substance D was formed to have molecular mass of 42,000 .Determine the number of molecules present in the substances(H+1 ,C=12) 2mks

- d. State
- The condition necessary for the conversion of ethanol to substance A.(1mk)
 - The catalyst required in the conversion of A and B.(1mk)

5. The table below gives the solubility of hydrated copper(ii) sulphate in mol dm⁻³ at different temperatures.

Temperature(°)	Solubility mol dm ⁻³
20	8 x 10 ⁻²
40	12 x 10 ⁻²
60	16 x 10 ⁻²
80	22 x 10 ⁻²
100	30 x 10 ⁻²

- On the grid provided plot a graph of solubility of copper(II) sulphate (vertical axis) against temperature.(3mks)*
- From the graph ,determine the mass of copper(II) sulphate deposited when the solution is cooled from 70°c to 40°. (Molar mass of hydrated copper(ii) sulphate = 250g)

b. In an experiment to determine the solubility of sodium chloride ,5.0 cm³ of a saturated solution of sodium chloride weighing 5.35g were placed in a volumetric flask and diluted to a total volume of 250cm³.

25.0 cm³ of the dilute solution of sodium chloride completely reacted with 24.1 cm³ of 0.1 M silver nitrate solution.



Calculate;

- i. Moles of silver nitrate in 24.1cm^3 of solution.(1mk)

- ii. Moles of sodium chloride in 25.0cm^3 of solution.(1mk)

- iii. Moles of sodium chloride in 250cm^3 of solution(1mk)

- iv. Mass of sodium chloride in 5.0cm^3 of saturated chloride solution (Na=23.0 Cu=35.5) (1mk)

- v. Mass of water in 5.0 cm^3 of saturated solution of sodium chloride(1mk)

- vi. The solubility of sodium chloride in g/100 g of water.(2mks)

6. The flow chart below shows some of the processes involved in large scale production of sulphur(vi) acid .

Use it to answer the questions that follow.

Sulphur(IV)oxide

oxygen

sulphur (vi) oxide

Oleum

Water

- a. Name the process

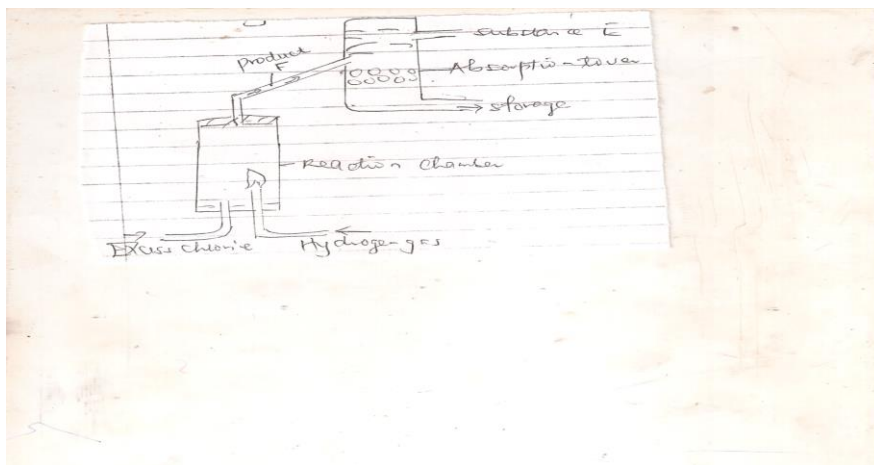
- b. i) Name substance A.(1mk)

ii) Write an equation for the process that takes place in the absorption tower.(1mk)

- c. Vanadium (v) oxide is commonly used catalyst in the process.

i. Name another catalyst which can be used for this process.(1mk)

- ii. Give two why reasons vanadium (v) oxide is commonly used catalyst.(2mks)
- d. State and explain the observations made when concentrated sulphuric (vi) acid is added to crystals copper(ii) sulphate in a beaker(2mks)
- e. The reaction of concentrated sulphuric (vi) acid with sodium chloride produces hydrogen chloride gas.State the property of concentrated sulphuric (vi) acid illustrated in the reaction.(1mk)
- f. Name two uses of sulphuric (vii) acid.2mks
7. The above diagram shows a set up that can be used for industrial manufacture of hydrochloric acid.Study it and answer the questions that follow.



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a. Name

i. Produce F

ii. Substance E

b. Explain an application of hydrochloric acid in the textile industry. (1mk)

c. Hydrochloric acid was added to iron powder in a test tube and shaken thoroughly to mix to 1cm^3 of the resulting solution, six drops of an aqueous solution of ammonia were added.

i. State the observation made on adding ammonia solution.

ii. Explain the observation stated above and write an ionic equation for the reaction. (2mks)

- iii. Concentrated hydrochloric is 35% pure with density 1.18g/cm^3 . Calculate its concentration in moles per litre..(3mks)